
Chapter 13 Properties Of Solutions Mater Academy Lakes

chapter 13 properties of solutions - colligative properties • changes in colligative properties depend only on the number of solute particles present, not on the identity of the solute particles. • how would the ionic compounds and covalent compounds behave..... **534 chapter 13 properties of solutions** - 536 chapter 13 properties of solutions in this instance, one of the resulting solutes is not ni metal but rather its salt NiCl_2 . if the solution is evaporated to dryness, $\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$ **chapter 13: properties of solutions - nclark** - 13.5 colligative properties colligative properties: properties depending on the number of solute particles in solution and not on the nature of the solute particles **chapter 13: properties of liquids - directory** - chapter 13: properties of liquids kahoot! 1. "like dissolves like" refers to similarities between ___ of two miscible solutions. molecular weights, intermolecular forces, shapes, densities 2. when NaOH dissolves, the container becomes hot. therefore, we can conclude that the magnitude of $\Delta H_{\text{solute}} + \Delta H_{\text{solvent}}$ ___ ΔH_{mixing} . >,